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Name: _____ Group: _____ Date: _____

Balancing Chemical Equations - Homework Sheet
Grade 10 Science

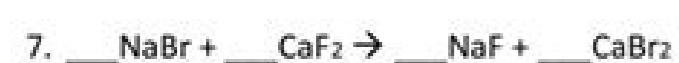
PART 1: Balance the following chemical equations

*Note, you may need to work out these balancing equations on extra paper

- $N_2 + H_2 \rightarrow NH_3$
- $S_8 + O_2 \rightarrow SO_3$
- $HgO \rightarrow Hg + O_2$
- $Zn + HCl \rightarrow ZnCl_2 + H_2$
- $SiCl_4 + H_2O \rightarrow H_4SiO_4 + HCl$
- $Na + H_2O \rightarrow NaOH + H_2$
- $H_3PO_4 \rightarrow H_4P_2O_7 + H_2O$
- $Si_2H_6 + O_2 \rightarrow SiO_2 + H_2O$
- $Al(OH)_3 + H_2SO_4 \rightarrow Al_2(SO_4)_3 + H_2O$
- $Fe + O_2 \rightarrow Fe_2O_3$
- $Fe_2(SO_4)_3 + KOH \rightarrow K_2SO_4 + Fe(OH)_3$
- $FeS_2 + O_2 \rightarrow Fe_2O_3 + SO_2$
- $Al + FeO \rightarrow Al_2O_3 + Fe$
- $Na_2CO_3 + HCl \rightarrow NaCl + H_2O + CO_2$
- $K + Br_2 \rightarrow KBr$
- $P_4 + O_2 \rightarrow P_2O_5$
- $C_2H_2 + O_2 \rightarrow CO_2 + H_2O$

Balancing Equations Problems

Name: _____ Date: _____ Pd: _____



Answers:

1. $1 N_2 + 3 H_2 \rightarrow 2 NH_3$
2. $1 S_8 + 12 O_2 \rightarrow 8 SO_3$
3. $1 HgO \rightarrow 1 Hg + 1/2 O_2$
4. $1 Zn + 2 HCl \rightarrow 1 ZnCl_2 + 1 H_2$
5. $1 SiCl_4 + 4 H_2O \rightarrow 1 H_4SiO_4 + 4 HCl$
6. $2 Na + 2 H_2O \rightarrow 2 NaOH + 1 H_2$
7. $2 H_3PO_4 \rightarrow 1 H_4P_2O_7 + 1 H_2O$
8. $1 Si_2H_6 + 3 O_2 \rightarrow 2 SiO_2 + 6 H_2O$
9. $2 Al(OH)_3 + 3 H_2SO_4 \rightarrow 1 Al_2(SO_4)_3 + 6 H_2O$
10. $4 Fe + 3 O_2 \rightarrow 2 Fe_2O_3$
11. $1 Fe_2(SO_4)_3 + 6 KOH \rightarrow 3 K_2SO_4 + 2 Fe(OH)_3$
12. $4 FeS_2 + 11 O_2 \rightarrow 2 Fe_2O_3 + 8 SO_2$
13. $2 Al + 3 FeO \rightarrow 1 Al_2O_3 + 3 Fe$
14. $1 Na_2CO_3 + 2 HCl \rightarrow 2 NaCl + 1 H_2O + 1 CO_2$
15. $2 K + Br_2 \rightarrow 2 KBr$
16. $1 P_4 + 5 O_2 \rightarrow 2 P_2O_5$
17. $2 C_2H_2 + 5 O_2 \rightarrow 4 CO_2 + 2 H_2O$

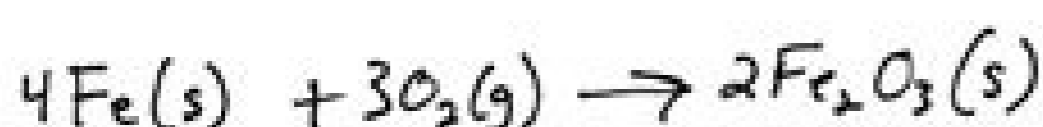
LOHS Honors Chemistry
Balancing Chemical Equations

Part I: Balance these chemical equations. Record what type of reaction on the blank in the right margin.

- $BaCl_2 + 2KClO_3 \rightarrow Ba(ClO_3)_2 + 2KCl$ DR
- $Pb(OH)_2 \rightarrow PbO + H_2O$ D
- $2Na + 2H_2O \rightarrow 2NaOH + H_2$ SR
- $Pb(NO_3)_2 + K_2CrO_4 \rightarrow PbCrO_4 + 2KNO_3$ DR
- $Al_2(SO_4)_3 + 3Ca(OH)_2 \rightarrow 2Al(OH)_3 + 3CaSO_4$ DR
- $Mg + 2HCl \rightarrow MgCl_2 + H_2$ SR
- $Sn + 2KOH \rightarrow K_2SnO_2 + H_2$ SR
- $2PaI_5 \rightarrow 2Pa + 5I_2$ D
- $3SiF_4 + 3H_2O \rightarrow 2H_2SiF_6 + H_2SiO_3$ SR
- $Cl_2O_7 + H_2O \rightarrow 2HClO_4$ S
- $16Na + S_8 \rightarrow 8Na_2S$ S
- $N_2O_3 + H_2O \rightarrow 2HNO_2$ S

Part II: Write balanced chemical equations for each of the reactions below. Assume that all ionic compounds are dissolved in water (aqueous). Also be sure include physical states where appropriate.

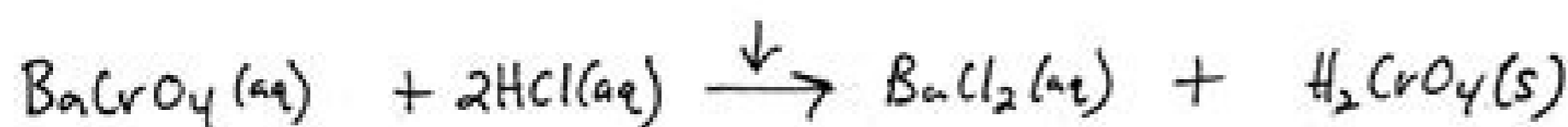
- Iron metal combines with oxygen gas to form iron (III) oxide solid



- Chromium metal and sulfuric acid (H₂SO₄) react to form chromium (III) sulfate solution and hydrogen gas.



- Barium chromate solution and aqueous hydrochloric acid (HCl) react to produce barium chloride solution and hydrogen chromate (solid).



Balancing Equations

- ___ H₂ + ___ O₂ => ___ H₂O
- ___ H₂PO₄ + ___ KOH => ___ K₂PO₄ + ___ H₂O
- ___ K₂ + ___ B₂O₃ => ___ K₂O + ___ B
- ___ HCl + ___ NaOH => ___ NaCl + ___ H₂O
- ___ Na + ___ NaNO₂ => ___ NaO + ___ N₂
- ___ C + ___ Si₂ => ___ CS₂
- ___ Na + ___ O₂ => ___ Na₂O
- ___ N₂ + ___ O₂ => ___ N₂O
- ___ H₂PO₄ + ___ Mg(OH)₂ => ___ Mg₃(PO₄)₂ + ___ H₂O
- ___ NaOH + ___ H₂CO₃ => ___ Na₂CO₃ + ___ H₂O
- ___ KOH + ___ HBr => ___ KBr + ___ H₂O
- ___ H₂ + ___ O₂ => ___ H₂O
- ___ Na + ___ O₂ => ___ Na₂O
- ___ Al(OH)₃ + ___ H₂CO₃ => ___ Al₂(CO₃)₃ + ___ H₂O
- ___ Al + ___ S₈ => ___ Al₂S₃
- ___ C₂H₆ + ___ N₂ => ___ C₂H₆N₂
- ___ Mg + ___ Cl₂ => ___ MgCl₂
- ___ Pb + ___ PbNO₃ => ___ Pb₂O + ___ N₂

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Live worksheets > English Finish!! Please allow access to the microphone Look at the top of your web browser. If you see a message asking for permission to access the microphone, please allow. Close This online quiz is intended to give you extra practice in balancing, identifying and predicting a random selection of over 150 chemical equations. Select your preferences below and click 'Start' to give it a try! Finish!! Please allow access to the microphone Look at the top of your web browser. If you see a message asking for permission to access the microphone, please allow. Close Welcome to It's Elemental - Balancing Act! The computer will give you a number of incomplete chemical equations. Balance the chemical equations by selecting coefficients from the pull-down menus. Once you think the equation is balanced, press the 'Check my answer!' button. Have fun and good luck! Other games based on the Table of Elements are available. A JavaScript enabled web browser is required. These scaffolded Balancing Chemical Equations Cornell Doodle Notes combine two effective note-taking strategies and can be used to introduce, teach, or review the concept of balancing chemical equations and the Law of Conservation of Mass. These notes review chemical formulas including coefficients and subscripts, differentiate between word, skeletal, and balanced equations, use two non-chemistry analogies to help students visualize what a balanced equation means, give steps for balancing an equation Page 2

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